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Q. Deep-sea divers use a breathing mixture of helium, oxygen, and carbon dioxide. What is the partial pressure of oxygen if the total pressure of the mixture is 760 mm Hg, the pressure exerted by the helium is 609 mm Hg, and the pressure exerted by the carbon dioxide is 1 mm Hg?

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Unit 11 • GPB 902 - Boyle's Law and Charles' Law <http://www.gpb.org/chemistry-study-of-matter/episodes/902> • GPB 903 - More About the Behavior of Gases

### GPB 902 - Boyle's Law and Charles' Law GPB 903 - More ...

Think about these steps to answer the question: Identify the system - P.E.). goes from liquid PE change, In this This is an type of change, the system (the water) heat from the surroundings. Identify the surroundings - This reaction is When Act E is high, the reaction is (slow, fast). Sketch a diagram of these reactions: thermic. faster, exothermic

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3 Problem Set #3  $I 1 = \text{_____}$   $I 2 = \text{_____}$   $I$

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Season 1 Episode 903 | 29m 59s More About the Behavior of Gases: The Combined Gas Law, Graham's Law of Diffusion, the Ideal Gas Law, Avogadro's Law, and Dalton's Law of Partial Pressure are ...

### Chemistry & Physics | Chemistry 903: More About the ...

Use Boyle's Law and Charles's Law to calculate and explain the relationship of pressure.

### Chemistry & Physics | Chemistry 902: Boyle's Law and ...

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