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Synthesizing Alum

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1211 Synthesis of Alum

Make Potassium Alum

(Potassium Aluminium

Sulfate) Video 5.

Experiment 5:

Preparation of alum:

vacuum filtration of a

solid Alum Synthesis

Lab Answers

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6. By adding heat to the solution, crystals were formed, which

decreased solubility. 7.

The crystals were small and white. They stuck together to form a little bit bigger white chunks.

Part B: 1. Yes, our melting temperature was 99.4 degrees C and the published melting temperature of alum is 92.5 degrees C.

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The Synthesis of Alum  
Lab by Michaela  
Tonsager

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Purpose-Synthesize a  
type of alum called  
potassium aluminum  
sulfate dodecahydrate-  
Observe and record the  
process of synthesizing,  
and calculate the percent  
yield of the synthesis



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Procedure 1. Wear goggles 2. Obtain a small piece of aluminum foil and measure its mass using the analytical balance. Next you ...

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[tensortom.com](http://tensortom.com)

The synthesis of alum proceeds in several reaction steps. The mole

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ratios of reactants and products can be found by combining the written equations for these separate reactions into an overall equation. Balance the overall equation for the synthesis of alum,  $\text{KAl}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O}$ , from aluminum, potassium hydroxide, sulfuric acid and water.

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## Synthesis of Alum from Aluminum

Mass of Alum yield =  
 $0.0397 \text{ moles} * 474.39$   
 $(\text{g} / \text{mole}) = 18.83 \text{ g}$   
yield =  $(\text{Actual yield (g)} / \text{Theoretical yield (g)}) * 100$   
 $= (14.14 \text{ g} / 18.83 \text{ g}) * 100 = 75.1 \%$

Discussion: The experimental value for the yielding of alum is 14.14 g and theoretical value is 18.83 g.

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Lab report on synthesis  
of Alum using  
Aluminum.

Synthesis of Alum:

$KAl(SO_4)_2 \cdot 12 H_2O$ . Chem  
111 Laboratory.

Synthesis of Alum:

$KAl(SO_4)_2 \cdot 12 H_2O$ .

Hazard Warning: The  
sodium hydroxide used  
in this experiment is  
highly corrosive. If you  
get it on your skin, wash

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immediately. If your skin feels slippery, that is a sign that you have gotten the sodium hydroxide on you.

Synthesis of Alum:

$KAl(SO_4)_2 \cdot 12 H_2O$

Synthesis of alum preliminary lab assignment answers.

Tuesday, May 3, AP  
Lab 1 □ Synthesis of  
Alum Posted by J.

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Using a hot plate, heat the acidified filtrate until all solid dissolves. Reaction steps involving the sulfuric acid: Only then begin to transfer the supernatant liquid from the beaker to the Buchner funnel.

Synthesis of alum  
preliminary lab  
assignment answers ...

The Synthesis of Alum  
*Page 14/26*

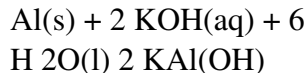
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Answers  
Overview of the  
Synthesis The synthesis  
of alum,  $KAl(SO_4)_2 \cdot 12H_2O$ , can be  
accomplished through  
the following reactions.

Aluminum is first  
oxidized by potassium  
hydroxide to form a  
soluble salt in the  
chemical reaction. 2



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Favourite answer Alum  
is aluminum sulfate.

Aluminum reacts with  
sulfuric acid to produce  
alum.

Chem Lab: Synthesis of  
Alum Question!?! |

Yahoo Answers

A synthesis reaction is a



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reaction in which two or more chemicals are combined to create a new compound or compounds. The following sequential reactions take place in this experiment to synthesize alum ( $\text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$ ): This synthesis demonstrates the aluminum hydroxide's ability to act

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The Formula, Synthesis,  
and Analysis of Alum -  
Odinity

When the solution is cooled, Alum crystals began to form. Finally, the Alum crystal was removed from the solution after 20 hours by filtration and washed with 12 v/v alcohol/water mixture.

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This wash liquid  
removes any  
contamination from the  
crystals but does not  
dissolve them.

Synthesis and Analysis  
of Potassium

Aluminium Sulphate ...

This is a video I made  
during my last lab.

Might get used in a lab  
report, not sure. The  
editing in this video is

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very basic. I just  
switched from Sony  
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Synthesis of Alum-

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Combine equations 1, 2, and 3 to find a full molecular equation for the synthesis of alum.

Notice that you will need to include

equations 2 and 3

multiplied by a factor of 2 to properly cancel the intermediate formed in

step 1. In a similar manner, combine

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equations 1i, 2i, and 3i  
to obtain a net ionic  
equation for the  
synthesis of alum.

CHEM 231 Experiment  
5 Synthesis of Alum -  
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Alum From Aluminum  
Lab Answers The  
Synthesis of Alum Lab  
Michaela Tonsager and  
Kaili Johnson  
Conclusion We

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determined that our sample was in fact alum. Our melting point of 99.4 degrees C was similar to the published melting point of 92.5 degrees C.

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Synthesis Lab Answers  
The Synthesis of Alum  
Lab Michaela Tonsager  
and Kaili Johnson

Conclusion We  
determined that our  
sample was in fact alum  
Our melting point of  
994 degrees C was  
similar to the published  
melting point of 925  
degrees C Our percent  
sulfate was 4244%,  
which is close



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reliefwatch.com

my ap chem lab analysis  
of alum the formula for  
epsom salts is  $\text{mgso}_4$   
 $7\text{h}_2\text{o}$  if 1250g of the  
compound is dissolved  
in water calculated the  
number of milliliters of  
200m  $\text{bano}_3$  which  
would be needed to just  
precipitate all of the

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sulfate as barium sulfate  
thanks for your help  
chemistry lab analysis  
of alum

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