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same atoms but at the
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them. These would result in a lower energy state than that of the atoms. They then enter a state called tunneling, where they stop exchanging energy with the atoms.

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*Response times vary
by subject and
question complexity.

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Median response time is 34 minutes and may be longer for new subjects. Q: For the reaction $2 \text{H}_2\text{O}_2(\text{l}) \rightarrow 2 \text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ $G^\circ = -227.8 \text{ kJ}$ and $H^\circ = -196.0 \text{ kJ}$ at 253 K and 1 atm. This... A: The balanced reaction taking place is given as ...

Answered: A

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formed when what?**

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packet, it is no
question easy then, in
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polyatomic ion. Your

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Answer should include
a discussion of
subatomic particles.

Somewhere in the ion
there are either extra
(in the case of an ion)
or fewer electrons (in
the case of a cation)
compared to the total
number of protons in
the atoms involved. 7.

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I am doing a packet
for my midterm and i
am very confused...

Someone please help
me answer these
questions regarding
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Ions? 2. How are the
atoms in a Polyatomic
ion put together? How
do you know? 3. How
does a Polyatomic ion
behave? Someone

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Answers will vary.

There are two extra electrons in the ion compared to a neutral S atom. There are two more electrons than protons in the S^{2-} ion. Two electrons have been added to an atom of sulfur. 6. The dotted line around

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each polyatomic ion in
Model 1 shows that
the group of atoms
has a. charge.

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have a -2 charge, and
each ion includes
both sulfur and
oxygen atoms. Sulfate
has four oxygen
atoms; sulfite only has
three. 9. The
"chlorate" polyatomic
ion has a charge of -1
and is composed of

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one chlorine atom
(the central atom) and
three oxygen atoms.

a.

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of the solution 02 g

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