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(d) Calculate the number of moles of precipitate that is produced in the experiment. $1.698 \text{ g} - 1.462 \text{ g} = 0.236 \text{ g}$
PbI₂ (s) 2 4 22 2 1 mol PbI 0.236 g

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Pbl = 5.12 10 mol Pbl

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