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Reaction in a Bag Air  
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Matters At Home  
Experiment

#1--Inflating an Airbag  
(baking soda and

vinegar) ~~Zero Waste~~  
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Chemical Reactions ~~The~~

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*Stoichiometry of a*

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Ward's Reaction In The  
Bag Inquiry  
Demonstration and Lab  
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Bags Of Reaction Lab  
Answers

reaction. Pre-lab  
discussion Read the  
entire laboratory  
investigation. Then

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answer the questions that follow. Pre-lab questions

1. Define reactants.
2. Define products.
3. How can you tell when a chemical reaction has happened?
4. What is the point of using a resealable bag?
5. What is the density of water?
- 6.



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Bags of reactions Lab - Mr.  
Niemann's Website

Bags of Reactions Lab

Introduction “Plop,  
plop, fizz, oh, what a  
relief it is,” claims an  
old television ad for a  
popular antacid. Just  
what is in the tablet that  
is relieving the upset  
stomach? What reaction  
is causing the fizzing?

Can you write a  
chemical equation for

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the process? With a bit of investigating, you will be able to discover answers to all these questions.

---

## 2. Bags of Reactions

Lab - Bags of Reactions

Lab Name ...

- The products of the reaction include sodium chloride ( $\text{NaCl}$ ), table salt; calcium carbonate

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( $\text{CaCO}_3$ ), the main component of chalk; and carbon dioxide ( $\text{CO}_2$ ), the metabolic “waste” gas exhaled during respiration. • The events that take place in the zipper-lock bag are part of a dynamic and complex reaction. Intermediate products may

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Reaction in a Bag -

Flinn Scientific

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answers 1/4

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This acid then converts some of the bicarbonate to carbon dioxide gas which begins to blow up the plastic bag.

Hydrogen ions +

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sodium bicarbonate --->  
carbon dioxide + water  
+ sodium ions H<sup>+</sup> ions  
+ NaHCO<sub>3</sub>---> CO<sub>2</sub> +  
H<sub>2</sub>O + Na<sup>+</sup> A good  
summary of the  
reactions (on Page 3) is  
at: <https://www.flinnsci.com/media/621024/91419.pdf>

---

Lab -- Reaction in a bag  
Download File PDF

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Reaction Lab Answers

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Answers Observe the

reaction until it comes  
to a complete stop.

Record your

observations. 5. When  
the reaction is complete,

record the mass of the  
bag and its contents in

Data Table 1. PART B

6. Add 2 teaspoons of

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calcium chloride to the  
second ...

---

Bags Of Reaction Lab  
Answers -

[vitaliti.integ.ro](http://vitaliti.integ.ro)

The lab is called Bags of Reactions. You put a tablet inside a resealable bag and record the reaction. The first reaction was a antacid tablet and water. The



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second reaction was  
 $\text{CaCl}_2$ ,  $\text{NaHCO}_3$ , and...

---

## CHEMISTRY

QUESTION - Yahoo

Answers

Heat the 50.0mL of HCl (aq) 70.0 °C. Place 0.200g of Mg (s) into heated HCl (aq). Use stopwatch to record the time of reaction from beginning to end.

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Perform steps 1-7 three times to obtain 3 trials. Perform steps 1-8 with HCl (aq) at 40.0 °C and 10.0 °C, keeping the amt of Mg (s) constant.

---

Rate of Reaction of HCl  
& Mg Lab Answers |  
SchoolWorkHelper  
Dispose of all the metals  
and burnt powders in a  
zip-lock bag and place

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in the trash. Clean and  
return all lab equipment.

IV. Data/Questions: Part

#1 Data Table . Sample

Appearance Before

Reaction Appearance

After Reaction 1.

Copper . 2. Magnesium.

3. Mercury (II) Oxide

---

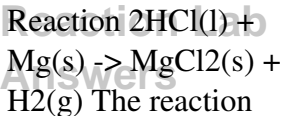
TYPES OF

CHEMICAL

REACTIONS LAB

*Page 19/33*

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between magnesium metal and hydrochloric acid is suitable for this particular experiment. The concentration of hydrochloric acid can easily and safely be adjusted by adding water to the solution.

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Concentration & Rate  
Factors Lab Answers |  
SchoolWorkHelper

- 1.) Measure 25 of water into a resealable plastic bag. Flatten the air out of the bag and seal it. Record the mass in Data Table 1.
- 2.) Record the mass of the antacid tablet in Data Table 1.
- 3.) Tip the bag sideways, and while holding the bag this

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Reaction Lab  
Answers

way, add the tablet so that the tablet and water do not mix. Do not trap any extra air in the bag.

---

Bags of Reactions - Mr.  
Wilkison's Science  
Website

Airbags have enough sodium azide to produce approximately 60L of gas in about 30 milliseconds.  $2\text{NaN}_3$ .

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In this experiment we will avoid the use of explosive reactions and concentrate on a slower gas-producing reaction.

---

## Stoichiometry Air Bag Lab Introduction

1. Put on your goggles and lab apron. Measure 25 mL of tap water into a re-sealable bag.

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Reaction Lab  
Answers

Flatten the air out of the bag and seal it. Record its mass in Data Table 1.

2. Record the mass of the antacid tablet in Data Table 1.

3. Tip the bag sideways, and while holding the bag this way, add the tablet so that the tablet and water do not mix. Do



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## MASS? Or The BAG OF REACTIONS LAB

Is this reaction in  
Equation 1

stoichiometrically  
balanced? \_\_\_\_\_

yes \_\_\_\_\_ Part 1 1. Fill  
the soda bottle with 1  
cup of vinegar. 2. Cut a  
small corner from the  
clear bag and add  $\frac{1}{4}$  tsp  
of baking soda into the  
bag fragment as shown  
below: 3. Carefully,

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drop the small bag into the soda bottle with the corner of the bag pointed

---

## Stoichiometry: Baking Soda and Vinegar Reactions

change in bag size, etc.

- The reaction takes about three to five minutes. During this time the student not

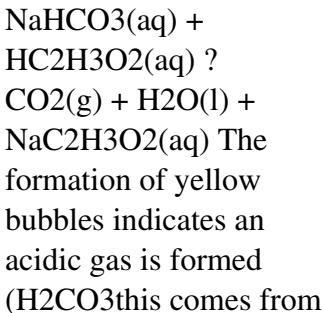
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holding the bag should write down any observations that the pair has made. NOTE: There is no danger of the bag exploding if the correct amounts of chemicals are used. Since everything is pre-measured, you should have no problems.

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and Chemistry in a  
Ziploc Bag ...

The first reaction in the bag is between baking soda and vinegar and is seen below:



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the interaction of  $\text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3$ ).

---

Reaction in a Bag  
( $\text{NaHCO}_3/\text{HC}_2\text{H}_3\text{O}_2$   
and  $\text{NaHCO}_3/\text{CaCl}_2$   
reaction rate - the  
change in concentration  
of a reactant or product  
over time. Teacher's  
Guide The Chemistry  
Matters teacher toolkit  
provides instructions

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and answer keys for  
labs, experiments, and  
assignments for all 12  
units of study.

---

Segment F: Air Bag Lab  
| Georgia Public  
Broadcasting

There are many  
different types of  
chemical  
reactions. There are  
single and double

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displacement reactions,  
combustion reactions,  
decomposition

reactions, and synthesis  
reactions. See if you can  
identify the type of  
reaction in this ten  
question chemical  
reaction classification  
practice test. Answers  
appear after the final  
question.

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Chemical Reaction  
Classification Practice  
Test

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