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The first step to balance

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the equation is to write down the chemical formula of reactants that are listed on the left side of the chemical equation. After this, you can list down the products on the right hand side of the chemical

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equation. There is an arrow between the sides, signaling the direction the reaction is happening in.

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...

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equations below: 1)  $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$  2)  $2 \text{ KClO}_3 \rightarrow 2$



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$\text{KCl} + 3 \text{O}_2 \rightarrow 2 \text{NaCl} + 1 \text{F}_2$   
 $\rightarrow 2 \text{NaF} + 1 \text{Cl}_2$  4)  $2 \text{H}_2 + 1$   
 $\text{O}_2 \rightarrow 2 \text{H}_2\text{O}$  5)  $1 \text{Pb(OH)}_2 + 2$   
 $\text{HCl} \rightarrow 2 \text{H}_2\text{O} + 1 \text{PbCl}_2$  6)  $2$   
 $\text{AlBr}_3 + 3 \text{K}_2\text{SO}_4 \rightarrow 6 \text{KBr} + 1$   
 $\text{Al}_2(\text{SO}_4)_3$  7)  $1 \text{CH}_4 + 2 \text{O}_2 \rightarrow$   
 $1 \text{CO}_2 + 2 \text{H}_2\text{O}$  8)  $1 \text{C}_3\text{H}_8 + 5$   
 $\text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$  9)  $2$

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$C_8H_{18} + 25 O_2 \rightarrow 16 CO_2 + 18 H_2O$   
10)  $1 FeCl_3 + 3 NaOH \rightarrow 1 Fe(OH)_3 + 3 NaCl$   
11)  $4 P + 5 O_2 \rightarrow 2 P_2O_5$

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To set up an equation in the  
Chemical Equations Gizmo,

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Answers Key  
type the chemical formulas into the text boxes of the Gizmo. First, type in "H<sub>2</sub>+O<sub>2</sub>" in the Reactants box and "H<sub>2</sub>O" in the Products box. This...

*Student Exploration-*

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A chemical equation is  
balanced by changing the  
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anything necessary DARK

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GREEN 10 What is the mass of  
Na?  $2\text{NaN}_3 \rightarrow 2\text{Na} + 3\text{N}_2$   
(500g) (?) (323g) 100 grams  
ORANGE 177 grams PINK 323  
grams YELLOW

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on the first Link or go to [https://phet.colorado.edu/en/  
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the "Run in HTML 5" BUTTON  
PRE-LAB: Use the simulation  
to adjust the coefficients

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formulae, it may not be balanced. The numbers of atoms of each element on the left must be the same as they are on the...

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Anne Marie Helmenstine,  
Ph.D. Updated January 04,  
2019. A balanced chemical  
equation gives the number  
and type of atoms  
participating in a reaction,

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Answers Key, products, and direction of the reaction. Balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow.



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The must find the equation that corresponds to what is inside each atom and then follow the arrow with the correct answer to the next

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atom. Once an atom is used  
it cannot be revisited, even  
if an arrow with the correct  
answer points back to it,  
the student must find  
another path and make there  
way to the end.

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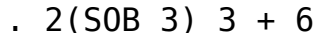
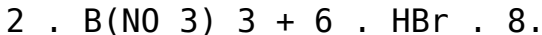
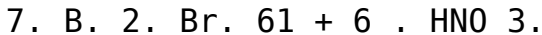
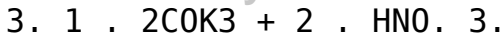
Balance the following  
chemical equations. 1. Fe .

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Answers Key

2 + 3 . H 2SO 4 . 1 . 2 (SOFe  
4) 3 + 3 . H . 2 . 2 . C . 2 . H .  
62 + 7 . 0 2 . 6 . 20 . H+ .  
4 . CO . 2 . 3 . KOH . 3 + 1 .  
H . 3PO 4 . 1 . 3PO 4K + 3 .  
H . 20 . 4 . SnO 21 + 2 . H 2 .  
1 . Sn + 2 . H 20 . 5 . NH 3 .  
4 + 5 . 0 . 2 . 4 . NO + 6 . H

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## *Answers* Key

Balancing Equations: Answers  
to Practice Problems. 1.

Balanced equations.

(Coefficients equal to one  
(1) do not need to be shown  
in your answers). (a)  $2 \text{Fe} + 3 \text{Cl}_2 \rightarrow 2 \text{FeCl}_3$ . (b)  $4 \text{Fe} + 3$

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Answers Key  
02  $\rightarrow$  2 Fe<sub>2</sub>O<sub>3</sub>. (c) 2 FeBr<sub>3</sub>  
+3 H<sub>2</sub>. SO<sub>4</sub>  $\rightarrow$  1 Fe<sub>2</sub>.

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There are 10 chemical  
equations for students to  
balance. Each equation comes

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Answers Key  
with a helpful chart so students practice counting up the number of atoms on each side and making sure they match. This worksheet is perfect for a student learning how to balance an equation for the first time!



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Give as classwork, assign as homework, or leave for a sub!

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