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1. Write the chemical and oxidation number of the positive ion. If the positive ion is monoatomic (only one atom), you can find its oxidation number from the periodic table. If the positive ion is polyatomic, use table 19.2 to find the oxidation number of the polyatomic ion.

19.2 Chemical Formulas

Flashcards

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3.2 Composition of Compounds A chemical formula is used to express the structure of a molecule. The formula tells which elements and how many of each element are present in a compound. Formulas are written using the elemental symbol of each atom and a subscript to denote the number of elements.

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Ionic bonds, covalent bonds and metallic bonds are examples of chemical bonds. The structure and bonding in a substance are modelled in different ways, including dot and cross diagrams.

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element valence electrons number
needed element valence electrons
number needed H11 Ne80 He20
Na17 Li17 Mg26 Be26 Al35 B35
Si44 C44 P53 N53 S62 O62 Cl71
F71 Ar80 Figure 19.3: Helium
atoms have only two electrons,
both of which are in the outermost
level. Helium is an exception to the
octet rule.

Changes in Matter Chapter 19 Molecules and

19.2 Chemical Formulas ions an
atom that has an electrical charge.
ionic compound A compound made
out of ions. Oxidation number
Shows how many electrons are

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