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pH of Acids,

Acids \u0026

Bases Tutorial pH

Calculations -

Calculate [H³⁰⁺]

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Calculating Ph
and $[OH^-]$, and
Find the pH of a
Solution

Calculating pH
& pOH,
[H⁺], [OH⁻],
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Bases CLEAR
& SIMPLE
pH, pOH, H₃O⁺,
OH⁻, K_w, K_a, K_b,
pK_a, and pK_b
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[H₃O⁺], [OH⁻] of
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pH and pOH

Calculations

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**Calculating the
pH of a Salt
Solution**

*Calculating pH
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*Concentration of
Hydronium K_a and
K_b Calculations*

*Solution,
Suspension and
Colloid |*

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pH of the

solution will be

governed by the

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following base

hydrolysis

equilibrium,

which is just

the reverse of

the

neutralization

reaction: $A^- +$

$H_2O \rightleftharpoons HA + OH^-$

K_b for A^- ,

which may be

calculated from

K_a of HA is K_b

$= K_w / K_a =$

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$$1.00 \times 10^{-14} / 1.00 \times 10^{-5} = 1.00 \times 10^{-9}$$

Making
the usual

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Pogil Answer Key*

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00+00:01

Subject:

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$\text{pH} = -\log[\text{H}^+] = -\log(1.0 \times 10^{-2}) = 2$

$\text{pOH} = 14 - \text{pH} = 14 - 2 = 12$

$[\text{OH}^-] = 10^{-\text{pOH}} = 10^{-12} \text{ M}$

$[\text{H}_2\text{O}] = 55.5 \text{ M}$

$\text{K}_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]}$

$1.0 \times 10^{-2} = \frac{(1.0 \times 10^{-2})[\text{A}^-]}{[\text{HA}]}$

$[\text{A}^-] = [\text{HA}]$

$[\text{A}^-] = 1.0 \times 10^{-2} \text{ M}$

$[\text{HA}] = 1.0 \times 10^{-2} \text{ M}$

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For Ap Chemistry

Strength Of

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Calculating pH -
Manasquan Public
Schools pOH =

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$$\begin{aligned} &= -\log(6.2 \times 10^{-5}) = 4.21 \text{ pH} \\ &= 14.00 - \text{pOH} = \\ &14.00 - 4.21 = \\ &9.79 \end{aligned}$$

6) +A
-solution with a
H concentration
of $1.00 \times 10^{-7} \text{ M}$
is said to be
neutral.

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schools poh

$= -\log[\text{oh}^-] =$

$-\log(6.2 \times 10^{-5})$

$= 4.21 \text{ ph} =$

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14.00 –

poh = 14.00

– 4.21 =

9.79 6) +a-

solution with a

h concentration

of 1.00×10^{-7} m

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Schools $pOH =$

$-\log[OH^-] =$

$-\log(6.2 \times 10^{-5})$

$= 4.21$ pH =

$14.00 - pOH =$

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14.00 - 4.21 =

9.79 6) +A

-solution with a

H concentration

of 1.00×10^7 M

is said to be

neutral. Page

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Keep in mind
that.

x .
is equal to the
concentration of
hydroxide ions
now in the
solution.

$[OH^-] = 6 \times 10^{-4} M$. Use this
value in the
equation for

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\displaystyle

$$\text{pOH} = -\log$$

$$[\text{OH}^-] = -\log$$

$$(6 \times 10^{-4}) = 3.2.$$

Remember that the sum of pH and pOH is always 14.

*Calculating pH
and pOH - High
School Chemistry*

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Schools pOH =

$-\log[\text{OH}^-] =$

$-\log(6.2 \times 10^{-5})$

$= 4.21 \text{ pH} =$

$14.00 - \text{pOH} =$

$14.00 - 4.21 =$

9.79 6) +A

-solution with a

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H concentration
of 1.00×10^{-7} M
is said to be
neutral.

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$$\begin{aligned} &= -\log[\text{oh}^-] = \\ &-\log(6.2 \times 10^{-5}) \end{aligned}$$

$$\begin{aligned} &= 4.21 \text{ ph} = \\ &14.00 \text{ \–} \end{aligned}$$

$$\begin{aligned} &\text{poh} = 14.00 \\ &\text{\–} 4.21 = \end{aligned}$$

9.79 6) +a-
solution with a
h concentration
of $1.00 \times 10^7 \text{ m}$

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