

## Chapter 11 Chemical Reactions

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Chemistry chapter 11 Chemical reactions answer key. coefficient. a whole number that appears before a formula in an equation. spectator ion. a particle not directly involved in a chemical reaction. combustion reaction. a reaction in which oxygen reacts with another substance, often producing light or heat. reactant.

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Chemical Reactions Chapter 11. precipitate. bubbles. signs of a chemical reaction. subscript. Solid compound produced from 2 aqueous solutions during a chem.... Gas given off during a chemical reaction; one of the signs of.... change in heat/light, formation of a gas (bubbles), solid prec....

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Write a balanced chemical equation for each reaction. Use the necessary symbols from Table 11.1 to describe the reaction completely. a. Bubbling chlorine gas through a solution of potassium iodide gives elemental iodine and a solution of potassium chloride. b. Bubbles of hydrogen gas and aqueous iron (III) chloride are produced when metallic

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Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride.  $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$  D

Chapter 11: Chemical Reactions

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Chapter 11: Chemical Reactions SC2 Students will relate how the Law of Conservation of Matter is used to determine chemical composition in compounds and chemical reactions. SC2.a. Identify and balance the following types of chemical equations: Synthesis, Decomposition, Single Replacement, Double Replacement, and Combustion SC2.b.

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cleavage reactions: A chemical reaction that has the general form  $\text{AB} \rightarrow \text{A} + \text{B}$ . Cleavage reactions are the reverse of condensation reactions.), and oxidation–reduction reactions A chemical reaction that exhibits a change in the oxidation states of one or more elements in the reactants that has the general form oxidant + reductant  $\rightarrow$  reduced oxidant + oxidized reductant.

Chapter 11.6: Types of Chemical Reactions - Chemistry ...

Chapter 11 Chemical Reactions 113 SECTION 11.1 DESCRIBING CHEMICAL REACTIONS.. A chemical reaction occurs when one or more \_\_\_\_\_ change into one Chapter a l to ChemiCal reaCtIons 4.1 Chemical Reactions and Chemical Equations 127 A chemical equation is a shorthand description of a chemical reaction.

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In the chemical reaction given, the reactants are sodium and water and the (formed) products are hydrogen and sodium hydroxide.

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As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium (III) oxide forms. We can describe this reaction with a chemical equationAn expression that gives the identities and quantities of the substances in a chemical reaction.

Chapter 11.3: Chemical Equations - Chemistry LibreTexts

A chemical reaction means a rearrangement of the configuration to acquire the electronic configuration of the nearest stable or noble gas by transferring or sharing the electrons. Question 5. What is a valence electron?

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Chapter 11 Chemical reactions What is a chemical reaction? its a reaction that changes the make up of an element and also creates a new substance. a chemical reaction occurs when compounds, elements, or atoms meet.

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