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### Reactions Answers

Chemistry chapter 11

Chemical reactions  
answer key.

coefficient. a whole number that appears before a formula in an equation. spectator ion. a particle not directly involved in a chemical reaction.

combustion reaction.

a reaction in which oxygen reacts with

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another substance,  
often producing light  
or heat. reactant.

Answers

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Section 11.1  
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Describe the steps in writing a balanced chemical equation.

Write the skeleton equation for the following reactions:

Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur

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dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride.  $CuS(s) + O_2(g) \rightarrow Cu(s) + SO_2(g)$  D

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Gravity. Created by.

tagreen348. Terms in  
this set (51) (aq)

dissolved in water (s)  
solid (l) liquid.

reactants. the starting  
substances of a  
chemical reaction.

products. the ending  
substances of a  
chemical reaction.

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Additional Questions

& Answers. Question

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1. Give an example of a complex chemical reaction. Answer:

‘Photosynthesis’ is a complex chemical reaction. Question 2.

Give an example of simple chemical reactions. Answer: Rusting of iron and combustion of fuels are simply chemical reactions. Question 3.

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Types of Chemical

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Reactions Often

charcoal briquettes provide the heat for barbeque grills

through the burning of carbon. Have you ever felt the heat and smelled the smoke coming from a burning charcoal grill?

The heat and smoke are the products of a combustion reaction.

Combustion is one of

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the 7ve general types  
of chemical

Reactions

### Answers

#### 11.2 Types of

#### Chemical Reactions

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Write correct formulas

for the products in

these double

replacement

reactions. 1)  $\text{H}_3\text{PO}_4$  2)

$\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow$  3)

$+ \text{Al}_2(\text{SO}_4)_3$  4)  $\text{AgNO}_3$

$+ \text{KCl}$  5)  $\text{H}_2\text{SO}_4 \rightarrow$

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+ 0+Co Q) ?BC2H302

3k! 0 + Poq +3CaS()

50 361) 10)

AgC2H302 + K2Cr04

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Chemical

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11.2 Types of

Chemical Reactions>

29 Whether one metal  
will displace another  
metal from a

compound depends  
upon the relative

reactivities of the 11.2

Types of Chemical

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### 11.1—Describing Chemical Reactions

The reactants are written on the left and the products on the right. The arrow that separates them is called yield. Chapter 11: Chemical Reactions

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### 11.2 Types of Chemical Reactions >

30 A halogen can also replace another halogen from a compound. •The activity of halogens decreases as you go down Group 7A of the periodic table—fluorine, chlorine, bromine, and iodine. •Bromine is

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more active than iodine, so this reaction occurs:  $\text{Br}_2(\text{aq}) + 2\text{NaI}(\text{aq}) \rightarrow 2\text{NaBr}(\text{aq}) + \text{I}_2(\text{aq})$

### 11.2 Types of Chemical Reactions>

Most chemical reactions can be classified into one or more of five basic types: acid–base

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Chemical Reactions  
Answers

reactions: A chemical reaction that has the general form  $AB + C \rightarrow AC + B$  or  $AB + CD \rightarrow AD + CB$ .,

condensation

reactions: A chemical reaction that has the general form  $A + B \rightarrow AB$ .

Condensation reactions are the reverse of cleavage reactions.

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As shown in Figure  
11.3.1, applying a  
small amount of heat

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to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium(III) oxide forms. We can describe this reaction with a chemical



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Chemical  
equation An

expression that gives  
the identities and  
quantities of the ...

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