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SECTION 11.1

DESCRIBING

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REACTIONS (pages
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explains how to write
equations describing
chemical reactions

using appropriate
symbols. It also

describes how to write
balanced chemical

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equations when given the names or formulas of the reactants and products in a chemical reaction.

Writing Chemical Equations (pages 321–323) 1.

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writing a balanced
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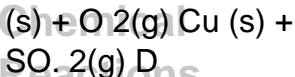
Write the skeleton
equation for the

following reactions:

Heating solid
copper(II)sulfide in the
presence of oxygen
gas produces pure
copper and sulfur
dioxide gas. Iron
metal and chlorine
gas react to form solid
iron(III)chloride. CuS

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PROBLEM 2 (page 324) 2. Sulfur burns in oxygen to form sulfur dioxide. Write a skeleton equation for this chemical reaction. Include appropriate symbols from table 11.1. Analyze Step 1. Identify the relevant concepts. Write the formula for each reactant and ...

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Reactions. In a

chemical reaction, the

reactants are written

on the left and the

products on the right.

The arrow that

separates them is

called yield.

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SECTION 11.1

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REACTIONS (pages
321–329). This

section explains

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PROBLEM 17 (page 334). 17. Complete the equations for these single-replacement reactions in aqueous solution. Balance each equation. Write “no reaction” if a ...

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reactions can be taken as well as picked to act.

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The initial step in solving a problem of this type must be to write the balanced chemical equation for the reaction.

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11.4.1 shows that it is balanced as written, so we can proceed to the strategy outlined in Figure 11.4.1, adapting it as follows:

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