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stoichiometry? 1. true

2. true 3. false 4. true

5. true 6. 2, 2, 64.10

7. 3, 3, 96.00 8. 2, 2,

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88.029. 4, 4, 72.08

10. methanol and
oxygen gas 11.

carbon dioxide and
water 12. 160.10 g

13. 160.10 g 14. They
are equal. 15. A mole
ratio is a ratio

between the numbers
of moles

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111.11.1 Objectives

Describe the types of relationships indicated by a balanced chemical equation.

State the mole ratios from a balanced chemical equation.

Review Vocabulary

reactant: the starting substance in a chemical reaction

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below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction

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In Section 11.3 , for
example, you learned
how to express the
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ammonium dichromate volcano in terms of the atoms,

ions, or molecules

involved and the

numbers of moles,

grams, and formula

units of each

(recognizing, for

instance, that 1 mol of

ammonium

dichromate produces

4 mol of water). This

section describes how

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to use the stoichiometry

stoichiometry of a reaction to answer

questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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