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Chapter 11 - Stoichiometry Section 11.1 What is stoichiometry?  
1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between the numbers

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of moles

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111.11.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry

*Chapter 11: Stoichiometry*

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Stoichiometry. stoichiometry. mole ratio.  
limiting reactant. excess reactant. the study  
of quantitative relationships between the  
amounts of... in a balanced equation, the  
ratio between the number of moles... a  
reactant that is totally consumed during a  
chemical reaction...

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• Chapter 11 209 StoichiometryStoichiometry

CHAPTER 11 SOLUTIONS MANUAL Section 11.1

Defining Stoichiometry pages 368–372 Practice

Problems pages 371–372 1. Interpret the

following balanced chemical equations in

terms of particles, moles, and mass. Show

that the law of conservation of mass is

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Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction

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In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to

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ensure complete combustion of a ...

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