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d. 400 cm 3 0.0004 m 2. For each measuring device shown below, identify the quantity measured and tell when it would remain constant ...

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Construct mole ratios from balanced chemical equations and apply these ratios in stoichiometric calculations. 12.2.2 Calculate stoichiometric quantities from balanced chemical equations using units of moles, mass, representative particles, and volumes of gases at STP. Guide for Reading

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Answers 1. a.  $2 \text{ Al}_2\text{O}_3 \rightarrow 4 \text{ Al} + 3 \text{ O}_2$  or  $2 \text{ Al}_2\text{O}_3 \rightarrow 4 \text{ Al} + 3 \text{ O}_2$  moles Al  $2 \text{ O}_3 = (2.3)(2)/4 = 1.2$  moles b. Using the same ratios, moles O  $2 = (3.9)(3)/2 = 5.6$  moles 2. a. 2 moles Fe gives 3 moles H  $2$ , moles H  $2 = (1.7)(3)/2 = 2.6$  moles b. 3 moles H  $2\text{SO}_4$  gives 1 mole product moles yield =  $3 \times 2.8 = 8.4$  moles 3. Mole ratios:  $2 \text{ mol Mg} / 2 \text{ mol MgO} = 1 \text{ mol Mg} : 1 \text{ mol product} : 1 \text{ mol O}$

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