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oxygen to produce aluminum oxide as follows: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$

a. If you use 2.3 moles of Al, how many moles of Al_2O_3 can you make?

b. If you want 3.9 moles of Al_2O_3 , how many moles of O_2 are needed?

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calculation of quantities
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Mole ratio. ... Use the
following balanced
equation to answer the
question: $\text{Mg} + 2\text{H}_2\text{O}$



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Stoichiometry . In the reaction represented by the equation $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$, how many grams of ...

Answer the questions above, assuming we started with 30 grams of ammonium nitrate and

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50 grams of sodium phosphate. Consider the following reaction:

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equations as a basis to
calculate how much
reactant is needed or
product is formed in a
reaction. Stoichiometry
= calculation of
quantities in chemical
reactions is a subject of
chemistry.

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placed the consequences
squarely in front of these
Hebrews. Those who
doubt Jesus have sinned
willfully, trampled under

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foot the Son of God,
regarded as unclean the
blood of Page 10/28

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and Attributions; You
are in charge of setting
out the lab equipment
for a chemistry
experiment. If you have
twenty students in the
lab (and they will be
working in teams of two)
and the experiment calls
for three beakers and
two ...

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pages 368 – 372

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371 – 372 1. Interpret

the following balanced

chemical equa-tions in

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terms of particles, moles,
and mass. Show that the
law of conservation of
mass is

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