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Chapter 12.1, 12.2 Stoichiometry p1 Unit
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Stoich Limiting Reactant

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Chapter 12 Stoichiometry 127. SECTION 12.1 THE ARITHMETIC OF

EQUATIONS (pages 353 – 358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles,

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representative particles, masses, and gas
volume at STP.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

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**SECTION 12.1 THE ARITHMETIC
OF EQUATIONS** (pages 353 – 358) This
section explains how to calculate the
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THE ARITHMETIC OF EQUATIONS
(pages 353 – 358) This section explains
how to calculate the amount of reactants
required or product formed in a

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nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

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**SECTION 12.1 THE ARITHMETIC
OF EQUATIONS** (pages 353 – 358) This
section explains how to calculate the
amount of reactants required or product
formed in a nonchemical process.

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Introduce the term stoichiometry in your own words. Stress that stoichiometry allows students to calculate the amounts of chemical substances involved in chemical reactions using information obtained from balanced chemical equations.

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12.1 The Arithmetic of Equations 12

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Chapter 3: Stoichiometry – Guided
Reading Section 3.1 – 3.2 1. True or
False? Most hydrogen atoms have a mass
of 1.008 amu. Justify your answer. If true,
explain why it is true. If false, what mass
do most hydrogen atoms have? False,
1.008 amu is actually hydrogen ' s average
mass, NO atom of hydrogen actually has

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