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using dimensional
analysis and/or
simple algebraic
formulas as
appropriate. SC2.d:
Identify and solve
different types of
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problems,
specifically relating
mass to moles and
mass to mass.

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the conceptual
principle of limiting
reactants.

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12.1 Stoichiometry

Intro. What is
stoichiometry?

Stoichiometry -

Defines the
quantitative
relationships

between amount of

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reactants used and
products formed.

Operates based on
Law of

Conservation of
Mass. Really its an
incredible
application of what
humans know about
matter in the 21st
century. We are
able to predict with
. extremely high
accuracy

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finally the miles are
converted to any
other unit of
measurement
related to the unit
mole as the problem
require.

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A In any
stoichiometry
problem, the first

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step is always to calculate the number of moles of each reactant present. In this case, we are given the mass of $K_2Cr_2O_7$ in 1 mL of solution, which we can use to calculate the number of moles of $K_2Cr_2O_7$ contained in 1 mL:

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step is always to
calculate the
number of moles

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in your tool box 1.

How many moles

CH_3OH are in

14.8 g CH_3OH ? 2.

What is the mass in

grams of 1.5×10

16 atoms S? 3. How

many molecules of

CO_2 are in 12.0 g

CO_2 ? 4. What is

the mass in grams

of 1 atom of Au?

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many moles CH₃

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different types of
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specifically relating
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Stoichiometry.
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Chapter 12.1 – The
Arithmetic of
Equations. A
balanced chemical
equation provides
quantitative
information.

Chemists use
balanced equations
as a basis to
calculate how much
reactant is needed
or product is
formed in a

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reaction. The calculation of quantities in chemical reactions is called stoichiometry.

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balanced equation,
the ratio between
the number of
moles.... a reactant
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consumed during a
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mol A mol A 1. How
many moles CH₃
OH are in 14.8 g CH
3 OH? 2. What is
the mass in grams
of 1.5×10^{16} atoms
S? 3. How many
molecules of CO₂
are in 12.0 g CO₂?
4. What is the mass
in grams of 1 atom
of Au? Tool Box:
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any stoichiometry
problem, the first
step is always to

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calculate the number of moles of each reactant present. In this case, we are given the mass of $K_2Cr_2O_7$ in 1 mL of solution, which we can use to calculate the number of moles of $K_2Cr_2O_7$...

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