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Nature of Covalent

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Problems 1. Draw

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electron dot structures for each molecule. a. chlorine b. bromine c. iodine What do you observe about the three structures? They all look similar with respect to the electrons. 2. The following molecules have single covalent bonds. Draw an electron dot structure for each. a. H_2O_2 b. PCl_3 3.

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62-67 Section 16.1 –
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Describe the difference between an ionic and a covalent bond. 24.

Classify the following compounds as ionic or covalent: a. MgCl_2 b. Na_2S c. H_2O d. H_2S 25. Explain why neon is monatomic but chlorine is diatomic. 26.

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to describe covalent
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When two atoms
combine in a bonding
situation, their atomic
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2. A δ bond is a
covalent bond in which
the electron density is
concentrated in the

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region along the internuclear axis; that is, a line between the nuclei would pass through the center of the overlap region.

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