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Chapter 16 - Covalent Bonding. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. nchen17. Terms in this set (25) Covalent Bonds - Form between nonmetals - Single, double, or triple bonds - Electrons are SHARED between atoms Example) CO₂, H₂O, Cl. Properties of Covalent Bonds

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Section 16.1 □ The Nature of Covalent Bonding Practice Problems 1. Draw electron dot structures for each molecule. a. chlorine b. bromine c. iodine What do you observe about the three structures? They all look similar with respect to the electrons. 2. The following molecules have single covalent bonds. Draw an electron dot structure for each. a. H₂O₂ b. PCl₃

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Chapter 16 Covalent Bonding Guided Reading Answers The sharing of electrons between atoms is called a covalent bond, and the two electrons that join atoms in a covalent bond are called a bonding pair of

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chemistry chapter 16 -- covalent bonding. covalent. 2 atoms in single covalent bond share h^{\square} . 2 atoms in double covalent bond share h^{\square} . 2 atoms in triple covalent bond share h^{\square} . nonmetals bond with each other by sharing electrons to become $^{\square}$. 2. chemistry chapter 16 covalent bonding Flashcards and Study ...

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16 Covalent Bonding Guided Reading Answers 16 covalent bonding Flashcards and Study ... Chapter 16: Covalent Bonding. Chemistry vocab and whatnot. STUDY. PLAY. Covalent bond. a chemical bond that involves sharing a pair of electrons between atoms in a molecule. Covalent compound. a chemical compound formed by the sharing of electrons. Molecule. two or more atoms Page 9/21 Chapter 16 Covalent Bonding Guided Reading Answers

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Polar Covalent Bonds. There are two types of covalent bonds: polar and nonpolar. In a polar covalent bond, shown in Figure 1, the electrons are unequally shared by the atoms and are attracted more to one nucleus than the other. Because of the unequal distribution of electrons between the atoms of different elements, a slightly positive (□+) or slightly negative (□□) charge develops.

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Chem chapter 8 lecture □ covalent bonding HW: read chapter 8.1 and make a vocab list
The covalent Bond-Why do atoms bond Atoms bond to gain stability within their electrons-In covalent substances, atoms share electrons to achieve this-Atoms are still trying to follow the octet rule o Apart from hydrogen which follows the duet rule (wants 2 valence electrons) What is a covalent bond?

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