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Chapter 17 Review Use the information below to calculate  $\Delta H$  for the reaction. Target equation:  $2\text{NO}_2(\text{g}) \rightleftharpoons \text{N}_2\text{O}_4(\text{g})$  Given equations:  $\text{N}_2(\text{g}) + 2\text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g}) \quad \Delta H = 67.7 \text{ kJ}$   $\text{N}_2(\text{g}) + 2\text{O}_2(\text{g}) \rightleftharpoons \text{N}_2\text{O}_4(\text{g}) \quad \Delta H = 9.7 \text{ kJ}$  13.

Chapter 17 Review How much heat is required to raise the temperature of  $6.5 \times 10^2 \text{ g}$  of aluminum by  $30 \text{ }^\circ\text{C}$ ?

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