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WORK Use the three-step problem-
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Consider 2 metals, A and B, each
having a mass of 100 g and an
initial temperature of 20 °C.

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Thermochemistry Practice
Problems (Ch. 6) 1. Consider 2
metals, A and B, each having a

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mass of 100 g and an initial temperature of $20\text{ }^{\circ}\text{C}$. The specific heat of A is larger than that of B. Under the same heating conditions, which metal would take longer to reach $21\text{ }^{\circ}\text{C}$?

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Problems (Ch. 6) 1. Consider 2 metals, A and B, each having a mass of 100 g and an initial temperature of $20\text{ }^{\circ}\text{C}$. The specific heat of A is larger than that of B. Under the same heating conditions, which metal would take longer to reach $21\text{ }^{\circ}\text{C}$? Explain

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17 Practice Problems In your notebook, solve the following problems. SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK Use the three-step problem-

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Problems
solving approach you learned in Chapter many kilojoules of energy are in a donut that contains 200.0 Calories? 2. What is the specific heat of a substance that has a mass of 25.0 g and requires

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Test Specific Heat Problems 1)

How much heat must be absorbed
by 375 grams of water to raise its
temperature by 25°C ? 2) What
mass of water can be heated from
 25.0°C to 50.0°C by the

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Problems addition of 2825 J? 3) What is the final temperature when 625 grams of water at 75.0°C loses $7.96 \times 10^4\text{ J}$? Chapter 17.

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A 12.2 g sample of an unknown metal sample was heated to $98.6\text{ }^{\circ}\text{C}$ and then put into a calorimeter containing 25.0 ml of water at $22.3\text{ }^{\circ}\text{C}$. After mixing, the temperature of the water and metal increased to a maximum of $28.4\text{ }^{\circ}\text{C}$ after 35 seconds.

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