

Chapter 19 Acids Bases Answer Key

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An acid is a substance that contains hydrogen and ionizes to produce hydrogen ions in aqueous solution; A base is a substance that contains a hydroxide group and dissociates to produce a hydroxide ion in aqueous solution. Bronsted-Lowry model. A model of acids and bases in which an acid is a hydrogen-ion donor and a base is a hydrogen-ion acceptor.

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Acid and Base Worksheet - Answers. 1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a) HNO₃ + OH⁻ (H₂O + NO₃⁻. HNO₃ and NO₃⁻ make one pair OH⁻ and H₂O make the other. b) CH₃NH₂ + H₂O (CH₃NH₃⁺ + OH⁻

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Prentice Hall Chemistry Chapter 19: Acids, Bases and Salts ... acids are hydrogen-containing compounds that ionize to yield hydrogen ions (H⁺) in aqueous solution. He also said that bases are compounds that ionize to yield hydroxide ions (OH⁻) in aqueous solution. Chemistry-Chapter 19-Acids, Bases, and Salts Flashcards... Chapter 19 Acids_ Bases_ and Salts - Free download as

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Jake Abrahams Chapter 19: Acids, Bases, and Salts Vocab Acid: a compound that produces hydrogen ions in a solution Base: a compound that produces hydroxide (OH⁻) ions in a solutionHydronium Ion: the positive ion formed when a water molecule gains a hydrogen ion Conjugate Acid: the particle formed when a base gains a hydrogen ion Conjugate Base: the