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Chapter 2: Introduction to the Chemistry of Life. Figure 2.1 Foods such as bread, fruit, and cheese are rich sources of biological macromolecules. The elements carbon, hydrogen, nitrogen, oxygen, sulfur, and

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phosphorus are the key building blocks of the chemicals found in living things. They form the carbohydrates, nucleic acids, proteins, and lipids (all of which will be defined later in this chapter) that are the fundamental molecular components of all organisms.

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Chapter 2 The Chemistry of Life What are the basic building blocks of all matter?

Atoms Describe the structure of an atom protons, electrons, and neutrons. The nucleus (center) of the atom contains the protons (positively charged) and the

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neutrons (no charge).

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Essential Question: What are the basic chemical principles that affect living things?

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Biology Chapter 2- The Chemistry of Life

Biology: Chapter 2, The Chemistry of Life.

Atom. Nucleus. Electron. Element. An atom is the smallest constituent unit of ordinary matter th.... The nucleus is the small, dense region consisting of protons a.... The electron is a subatomic particle, symbol e^- or e^{-} , with a....

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Chapter 2 the chemistry of life Flashcards
and Study Sets ...

A B; atom: the basic unit of matter: nucleus:
the center of the atom: electron: a negatively
charged particle: element: a pure substance
that consists entirely of one type of atom

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Quia - Chapter 2: The Chemistry of Life Vocabulary Review

Chemistry 1405 Chapter 2 1. chemistry.

Chemistry is the study of matter, its properties, how and why substances combine or separate to form other substances, and how substances interact with energy. 2. matter. The term matter refers to

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anything that occupies space and has mass—in other words, the “ stuff ” that the universe is made of. 3. three different levels of matter – The three ...

1405 - Chapter 2 Vocab(2) (1).rtf -
Chemistry 1405 Chapter ...

A compound is a substance formed by the

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Key
chemical combination of two or more elements in definite proportions. (this means that H_2O is water, but H_3O is not).
Compounds have different physical and chemical properties from the elements they are made of

Biology: Chapter 2, The Chemistry of Life

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Flashcards

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hard glass tube about 60 cm in length. It is sealed at both the ends and ...

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Chapter 2 The Chemistry of Life What do you see when you look at this picture? Is it just a mass of tangled ribbons? Look closely. It ' s actually a complex pattern of three-dimensional shapes. It represents the structure of a common chemical found

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inside living cells. The chemical is a protein called kinase.

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Nonmetals; 18.5 Occurrence, Preparation,
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Occurrence, Preparation, and Properties of

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Carbonates; 18.7 Occurrence, Preparation,
and Properties of Nitrogen

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2. What relationship exists between the mass number of an element and isotopes of that element? 3. Explain the difference between

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Key ionic and covalent bonds. 4. Compare and contrast adhesion and cohesion. Below is a guide for Chapter 2 Test

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that make up the at.... The atomic number.
A chemical reaction in which two molecules
are bonded together.... A chemical process
that splits a molecule by adding water.
Protons, Electrons and Neutrons. The
number of protons.

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Study Sets ...

This chapter looks at atoms, bonds, pH and organic molecules. Good review of chemistry we see in microbiology.

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Key Concepts Class Date Section Review

2-4 Completion On the lines provided,
complete the following sentences. 1.

Chemical reactions that energy often occur
spontaneously. 2. During a chemical
reaction, chemical bonds are 3. Biological
catalysts, or enzymes, act by lowering the
required for a reaction. 4.

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