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substances which have an unchanging chemical composition all over them. For example, Water, gold, Pure metals are some of pure metals.

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container has volume of , and holds steam at C. 1/4 of the volume is in liquid point and the remaining at vapor form. Determine the pressure of the steam, and quality of the saturated mixture, and density of the mixture. Given:

Volume (V)

Temperature (T) C.

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Substances Phase

Change and Property
Diagrams In this

chapter we consider
the property values
and relationships of a
pure substance (such
as water) which can

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exist in three phases
– solid, liquid and
gas.

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Substances –
Thermodynamics

The change in
internal energy can
be found from the
first law of

thermodynamics: U
 $= Q - W = (3.5 \times 10$
 $J) - (2.1 \times 10 \quad J) =$

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0.9 x 10³ J = 90 kJ.)

A gas in a cylinder is kept at a constant pressure of 3.5×10^5 Pa while 300 kJ of heat are added to it, causing the gas to expand from 0.9 m^3 to 1.5 m^3 .

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3-22 is reconsidered.

The missing properties of water are to be determined using EES, and the solution is to be repeated for refrigerant-134a, refrigerant-22, and ammonia.

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-Thermodynamics is the study of the relationship between the energy transformation in the system and other physical quantities such as temperature, pressure and volume (P, V, T). -A thermodynamic equation of state is a

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mathematical relationship of the thermodynamic or state variables, such as pressure, volume and temperature.

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transfer is energy in
transition due to a
temperature
difference. The three
modes of heat
transfer are
conduction,
convection, and
radiation. Conduction

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Conduction heat transfer is a

progressive exchange

of energy between

the molecules of a

substance. Fourier's

law of heat

conduction is $Q = -kA \frac{dT}{dx}$

where Q !

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