Chapter 8 Covalent Bonding Work Answers

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Section 8.4 – Polar Bonds and Molecules. Covalent bonds involve sharing electrons between atoms. When the atoms in the bond pull equally, the bonding electrons are shared equally, and the bond is nonpolar. When the atoms in the bond pull unequally, the bonding electrons are pulled closer to one atom, and the bond is polar.

Chapter 8 – Covalent Bonding
Chapter 8: Covalent Bonding. Matter takes many forms in nature: In this chapter, we are going to learn to distinguish the type of compound that we have already studied, the “ioniccompound” (which contains oppositely-charged particles: metal cationsand non-metal anions), from a different type of compound – a “molecularcompound”.

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When H forms a bond with H O to form the hydronium ion H O , this bond is called a coordinate covalent bond because a. both bonding electrons come from the oxygen atom. b. it forms an especially strong bond. c. the electrons are equally shared. d. the oxygen no longer has eight valence electrons.

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covalent bonds. The majority of covalent bonds form between nonmetallic elements. 8. Describe how the octet rule applies to covalent bonds. Atoms share valence electrons; the shared electrons complete the octet of each atom. 9. Illustrate the formation of single, double, and triple covalent bonds using Lewis structures.

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242 Chapter 8 • Covalent Bonding Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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Covalent Bonding - Chapter 8. 1. Covalent Bonding Or How I Learned to Love Sharing (But Remember, File Sharing is Illegal) 2. As you should remember, ionic compounds are solids at room temperatures that have one ion strip the electron (s) from the other elements’ electron cloud.

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