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of elements in  
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mass relationships  
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left. Show all your work in the space provided. 1.

The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(\text{s}) \rightarrow$

$2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many moles of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed? 2.

Given the following equation:  $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow$   
 $( 2\text{HF}(\text{g})$

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### CHAPTER 9 REVIEW

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Section 1 Introduction  
to Stoichiometry

Objective • Define  
stoichiometry. •

Describe the importance  
of the mole ratio in  
stoichiometric  
calculations. • Write a  
mole ratio relating two  
substances in a chemical  
equation.

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calculations involving  
the mass relationships of  
elements in compounds.

Reaction Stoichiometry.  
calculations involving  
the mass relationships

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between reactants and products in a chemical reaction . Mole Ratio ...

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SECTION 3

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1.

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88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the following equation:

$$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$$

$N_2$ ; 2.0 mol a. Which chemical is ...

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3 Limiting Reactants

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molar relationships to



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determine the amounts  
of unknown reactants or  
products from the  
amounts of known  
reactants or products.

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Search for: 9.3 Solution Stoichiometry. Learning Objectives. By the end of this section, you will be able to: Perform stoichiometric calculations involving solution molarity; As we have seen in lab, many reactions such as single or double displacement reactions are carried out in aqueous medium (i.e. in water). Because these

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reactions occur in  
aqueous solution, we  
can use ...

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STOICHIOMETRY.

What is Sto-Key-Om-It-

Tree? Stoichiometry is

derived from the Greek

words stoicheion

“ element ” and

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metron “measure”.

Stoichiometry is defined as the branch of chemistry that deals

with the mass

relationships between

elements in a compound

and the mass

relationships between

reactants and products

in a chemical reaction.

There are 2 main types

of stoichiometry ...

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11.2. As you read this

section, summarize each

step on a tab and

include an example of

the step. 368 Chapter

11 • Stoichiometry

Section 11.11.1

Objectives Describe the

types of relationships

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indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a ...

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