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1. Chemical
equations and
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Objectives:
 List three observations that suggest that a chemical reaction has taken place.
 List three requirements for a correctly written chemical equation.
 Write a word equation and a formula equation for

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a given chemical
reaction.
 Balance a
Chemical
formula equation by
inspection.
 .

Answers

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than before future.

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admission it
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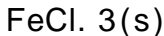
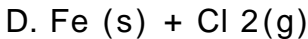
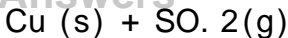
Answers Review

Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur

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dioxide gas. Iron
metal and chlorine
gas react to form
solid

iron(III)chloride.



Chapter 11:

Chemical Reactions

11.1 Describing

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Worksheet Review

Answers together
with Chapter 3

Chemical Reactions
and Reaction

Stoichiometry Pp.

Remember that

these answers can
be found in

different kinds of
factors. You may be
looking for

examples, ratios,
weight, sizes, and

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so on. Describing Chapter 11.1

Describing

Chemical Reactions

A representation of a chemical reaction using the chemical formulas of the reactants and products; a balanced chemical equation contains equal numbers of

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atoms of each
element on both
sides of the
equation. Four
indications that a
chemical reaction
has taken place. 1)
The release of heat
and/or light. 2)
production of gas.

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Describing
Chemical Reactions

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0132525763,

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Chapter Answers 1.

The A correctly written chemical equation must represent the known facts, contain the correct formulas for the reactants and products, and satisfy the law of conservation of mass. a.

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Chapter 7 Chemical
Reactions Section 1
Review Answers

CHAPTER 8

REVIEW Chemical
Equations and
Reactions SECTION
2 SHORT ANSWER

Answer the
following questions
in the space
provided. 1. Match
the equation type
on the left to its

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the right. c

synthesis (a) AX

BY AY BX d

decomposition (b)

A BX AX B b

single-displacement

(c) A B AX a

double-

displacement (d)

AX A X

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tools. Search. ... a
starting substance
in a chemical
reaction. ...
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reactions, show

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reactants on the left
and products on the
right, ...

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Chemical Equations
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Describing
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I. Introduction A.
Reactants 1.

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Original substances
entering into a
chemical rxn B.
Products 1. The
resulting
substances from a
chemical rxn
Reactants A
Products C.
Chemical Equation
1. Represents with
symbols and
formulas, the
identities and

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relative amounts
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Chapter review
Answers
answer key 1.

Chapter Review

Answer Key (Unit
5) 1. Energy in the
form of sound, light
or heat (absorbed
or released),
production of a gas,

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formation of a precipitate, and a color change 3. A coefficient is a number that appears in front of a formula in a chemical equation.
7.

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answer key -

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1 SHORT ANSWER
1. a. d b. a c. b d. f
e. e f. c 2. 8,4,9 3.
a. 12 atoms b. 16
atoms c. 51 atoms

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d. 3 1024 atoms

CHAPTER 8

REVIEW Chemical Equations and Reactions

Key Concept

Review 1. Two types of chemical reactions discussed in this section are synthesis reactions and decomposition reactions. 2. The

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products of a neutralization reaction are a salt (a compound consisting of a positive metal ion or polyatomic ion and a non-metal ion or polyatomic ion) and water. 3. (a) synthesis chemical reaction: $A + B : AB$

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Instruction 2.

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Quizzma A chemical
change or chemical
reaction is a
process in which
one or more pure

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substances are
converted into one
or more different
pure substances.

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