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The given base is OH^- . OH^- accepts H^+ H_2O Now, add a proton to the given base to get its conjugate... The given chemical equation for the formation of hydrogen iodide is shown below. $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightarrow 2\text{HI}(\text{g})$... Polarizability of atoms tends to decrease from left to right across a period and increases down a...

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In the first, the end nitrogen atom is 1, the central N atom is +1, and O atom is 0. For the second structure, the end N atom is 0, the central N atom is still +1, but the O atom is 1. The second is preferred because it places the negative formal charge on the most electronegative atom.

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2A.28 The coulombic attraction is inversely proportional to the distance between the two oppositely charged ions (Equation 1) so the ions with the shorter radii will give the greater coulombic attraction. The answer is therefore (c) Mg^{2+} , O^{2-} . 2A.30 The Br^{-} ion is smaller than the I^{-} ion (196 vs 220 pm).

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