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Lecture Notes $0.79 \text{ Mnitrogen} + 0.21 \text{ Moxygen} =$

$0.79(28) + 0.21(32) = 28.84$ The mass

fractions of oxygen and nitrogen in air are then $m_{\text{oxygen}} = (0.21)(32)/28.84 = 0.233$, or 23.3% and $m_{\text{nitrogen}} =$

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4. Write additional data required to solve the problem and the chemical equations if the process involves chemical reaction. 5. Select a suitable basis of calculations. 6. List by symbols each of the unknown values of the stream flows and compositions 7. Make a number of independent material balances equations equal to unknown

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view the study of the field of chemical engineering as a tree with material and energy balances being the trunk and the subjects of thermodynamics, fluid flow, heat transfer, mass transfer, reactor kinetics, process control, and process design being the branches off the trunk. From this perspective, it is

Basic Principles and Calculations in Chemical Engineering

$z + z_P$. $z = \hat{g}$ (2.2.8) Setting the limit on the left hand side as $z \rightarrow 0$ gives the definition of the derivative of P with respect to z : $\frac{dP}{dz} = -\hat{g}$ (2.2.9) In other words, the change in pressure, P , with height, z , is negative - pressure reduces with height (or increases with depth - as any diver knows).

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This page provides all lecture notes for the MIT course 10.34 Numerical Methods Applied to Chemical Engineering of Fall 2015, taught by Prof. William Green, Jr. and Prof. James W. Swan.

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