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ionic: $\text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$ net ionic: $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ 3. $\text{H}_2\text{SO}_4(\text{aq}) + \text{Mg}(\text{OH})_2(\text{aq}) \rightarrow \text{MgSO}_4(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$ ionic: $2\text{H}^+(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + \text{Mg}^{2+}(\text{aq}) + 2\text{OH}^-(\text{aq}) \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$ net ionic: $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ 4. $\text{NH}_4\text{OH}(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow \text{NH}_4\text{Cl}(\text{aq}) + \text{H}_2\text{O}(\text{l})$ ionic: $\text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{NH}_4^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$

9-1 GCSE Chemistry Ionic Equations Answers

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Experimental Errors, Significant Figures, and Net Ionic Equations Lab Section Name Date 5. Iron III nitrate and potassium sulfide I 6. strontium acetate and phosphoric acid - 7. aluminum sulfate and ammonium sulfide 8. barium hydroxide and zinc(II)

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$2\text{K}^+ (\text{aq}) + 2\text{Cl}^- (\text{aq}) + \text{Pb}^{2+} (\text{aq}) + 2\text{NO}_3^- (\text{aq}) \rightarrow 2\text{K}^+ (\text{aq}) + 2\text{NO}_3^- (\text{aq}) + \text{PbCl}_2 (\text{s})$ Net Ionic Equation: $2\text{Cl}^- (\text{aq}) + \text{Pb}^{2+} (\text{aq}) \rightarrow \text{PbCl}_2 (\text{s})$ Directions: Write balanced molecular, ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in aqueous solution.

Net Ionic Equation Worksheet Answers

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Chemistry 1a Net Ionic Answer Answer to 1a. Identify whether each species functions as a Brønsted-Lowry acid or a Brønsted-Lowry base in this net ionic equati... Solved: 1a. Identify Whether Each Species Functions As A B ... $2\text{K}^+ (\text{aq}) + 2\text{Cl}^- (\text{aq}) + \text{Pb}^{2+} (\text{aq}) + 2\text{NO}_3^- (\text{aq}) \rightarrow 2\text{K}^+ (\text{aq}) + 2\text{NO}_3^- (\text{aq}) + \text{PbCl}_2 (\text{s})$ Net Ionic

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occurs when barium sulfite (s) and excess nitric acid (aq) are combined. Note: Sulfites follow the same solubility trends as sulfates.

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