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A chemistry student has 4 empty cubes. Two of the cubes have 1 cm sides and masses of 0.5 g. The other two cubes have 2 cm sides and masses of 1.0 g. The student adds 1 or more balls to each cube, as shown in the table.

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Unit 6 • Reaction signs

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and symbols (+, , ,  
etc.) • Writing balanced  
equations o From

skeleton equations o By  
listing coefficients o

From reaction statements

Chemical formulas from  
names (nomenclature)

- Reaction

Classification System 1 o

Synthesis o Combustion

o Decomposition o

Single replacement o

Double replacement Unit



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7 • Oxidation-  
reduction (redox)  
reactions Oxidation

states/numbers Oxidized  
and reduced elements  
Activity series of ...

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not the test. This is a

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guide to the type of  
question that will appear  
on the test. Page 1 of 5 #  
AL COS # 1. #4.0 What  
type of solution that  
contains all of the solute  
it can hold at a given  
temperature? 2.

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CHEM 1210 Practice

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Exam #1 The following practice exams were given as a first exam in Chemistry 121 in previous quarters at Ohio State. While these exams serve as a guide for you to study, you should not take them literally and should choose questions from these exams that correlate to the learning objectives we went over this semester in class ...

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## Chemistry Final Exam Study Guide Section 1

Learning chemistry is about the understanding of the key concepts and the application of these concepts onto problem solving in chemistry.

These two key ingredients are also the focus of this course.

Unlike traditional textbooks, this Rapid

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