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Name _____

CHEMISTRY: A

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KEY I. For each of the following statements, write [I] for ionic, [C] for covalent

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each of the following

statements, write I

for ionic, C for

covalent

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Spring 2012 ... List
the appropriate SI unit
for the following
measurements.

Define temperature
and mass. a. Length
b. mass c. volume d.
temperature e. time f.
quantity 9. Perform
the following unit
conversions: Use the
factor label

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(dimensional analysis)
method.

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Chemistry Unit 5

Review Sheet

Answers

The AP Chemistry course is taught as a continuation of the first year course. As such, units are numbered accordingly. Units 1-8 are covered during the first year course and reviewed over a short time span during

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the AP course. (Links for these units can be found under the "CHEMISTRY" tab above and on the links immediately below this overview.)

Units 1-5 Pre-AP
Chemistry Review -
Mrs. Forest's ...
Chemistry Unit
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Read the descriptions

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Unit 3 Answers

below and determine whether each item describes an endothermic or an exothermic reaction.

- _____ 1. A reaction that produces heat energy _____ 2. A reaction that absorbs energy _____ 3.
Paper burns _____ 4

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Unit 5 Review. Show
all questions <= ... 2.5

moles per second;

Given the following

reaction: $C_3H_8(g) +$

$5O_2(g) \rightarrow 3CO_2 + 4$

$H_2O(g)$ If the initial

rate of the reaction

produces 2 moles of

water vapor, H_2O ,

per second, what is

the rate at which

propane, C_3H_8 , is

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consumed initially? ?

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36.5 g H M = moles of
solute liters of solution
0.200 M = mol solute

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0.250 L mol solute =
0.0500 mol HCl? g
HCl = 0.0500 mol HCl !
Cl 1 mol HCl = Cl 4. A
solution of Na₂CO₃
contains 65.0 g of
solute dissolved in
water to make a 3.00
M solution. What is
the volume of the
solution, in liters? ? g
HCl = 0.0500 mol HCl
! 36.5 g HCl 1 mol
HCl = 1.83 ...

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10-24a,b,c,d Unit 10
Review Sheet-Key

Chemistry 12 Unit 1 -
Reaction Kinetics Unit
1 - Review Sheet

Page 2 2. For each of
the following reactions
find a quantity or
property which could
be monitored in order
to measure the rate of
reaction See p. 2-5 in
SW. ("a" is done as

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an example.) a)
 $3\text{H}_2(\text{g}) + \text{N}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ pressure will
decrease as reaction
proceeds because

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