

## Chemquest 53 Answers

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PLEASE NOTE: If you have a question about these answers, it is your responsibility to come to office hours or ask during class work time. UNIT 12 - HW Practice Keys - ChemActivity 51: Cell Voltage - ChemActivity 50: Electrochemical Cell UNIT 11 - HW Practice Keys - ChemQuest 55: Free Energy - ChemQuest 54: 2nd Law of Thermodynamics

*HW Keys - Roosevelt High School AP Chemistry 2017-18*

This video is designed to help you complete the molarity chemquest found in Unit 11.

*Molarity*

118 ChemQuest 39 Name: \_\_\_\_\_ Date: \_\_\_\_\_ Hour: \_\_\_\_\_ Information: Molarity Concentration is a term that describes the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions contain only a little.

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An element exists as 4 different isotopes. 4.35% have a mass of 49.9461, 83.79% have a mass of 51.9405, 9.50% have a mass of 52.9407 amu, and 2.36% have a mass of 53.9389 amu. a. What is the average atomic mass for this element? b. What is the identity of this element?

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15.Given your answer to question 14 and the reaction you wrote in question 13, calculate the molarity of the nitrate ions (NO<sub>3</sub><sup>-</sup>) in the solution.

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(Base the answer to this question on the number of moles of propane that actually get combusted—which is your answer to part a.) 12 moles. For every mole of propane that combusts 3 moles of CO<sub>2</sub> are produced, so the number of moles of CO<sub>2</sub> that can be produced when 4 moles of propane combusts = 4(3) = 12.

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