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To calculate the answer here, we need to recall the formula: $\text{density} = \frac{\text{mass}}{\text{volume}}$ In this

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Answers, the mass is 2460 kg and the volume is 1.2 m^3 , so we simply need to substitute these values into the formula as follows: $\text{Density} = 2460 \text{ kg} \div 1.2 \text{ m}^3 = 2050 \text{ kg/m}^3$

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Answer Key For Density Worksheets - Teacher Worksheets

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Answers 1) $d = 40.5 \text{ g} / 15.0 \text{ mL}$ $d = 2.7 \text{ g/mL}$ 2) $d = 305.4 \text{ g} / 22.5 \text{ mL}$ $d = 13.57333 \text{ g/mL}$ 3) $d = \text{g} / \text{mL}$ $\text{g} = (d) (\text{mL})$ $d = 2.70 \text{ g/mL}$ $d = 13.6 \text{ g/mL}$ $\text{g} = (0.789 \text{ g/mL}) (200.0 \text{ mL}) = 158 \text{ g}$ 4) $(8.4 \text{ cm}) (5.5 \text{ cm}) (4.6 \text{ cm}) = 212.52 \text{ cm}^3$ $d = 1896 \text{ g} / 212.52 \text{ cm}^3 = 8.9 \text{ g/cm}^3$ Significant figures in the answer

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Answers are dictated by the length
measurements of two sig figs. 5) mass
of $\text{CCl}_4 = 703.55 \text{ g} - 345.8 \text{ g} = 357.75$
 g $d = 357.75 \text{ g} / 225 \text{ mL} = 1.59 \text{ g/mL}$
6) $d = 65.14 \text{ g} / 35.4 \text{ mL}$ $d = 1.84 \text{ g/mL}$
7) $d = \text{g} / \text{mL}$ $\text{g} = (d) (\text{mL} \dots$

density practice problems key -

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Density Practice: Worksheet #1.

Calculate density, and identify substances using a density chart.

Density is a measure of the amount of mass in a certain volume. This physical property is often used to identify and classify substances. It is

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Answers expressed in. grams per cubic centimeters, or g/cm^3 .

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1) A block of aluminum occupies a volume of 15.0 mL and weighs 40.5 g. What is its density? 2) Mercury metal is poured into a graduated cylinder

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Answers that holds exactly 22.5 mL. The mercury used to fill the cylinder weighs 306.0 g. From this information, calculate the density of mercury.

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Answers
1) A block of aluminum occupies a volume of 15.0 mL and weighs 40.5 g. What is its density? 2) Mercury metal is poured into a graduated cylinder that holds exactly 22.5

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Answers One way to do this is to mentally compare your answer against the density of water (1 gram per cubic centimeter). Light substances would float on water, so their density should be less than that of water. Heavy materials should have density values greater than that of water.

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$m(\text{silver}) = D \times v$ $m(\text{gold}) = D \times v$

$m(\text{silver}) = 10.5 \text{ g/cm}^3 \times 5 \text{ cm}^3$ $m(\text{gold})$

$= 19.3 \text{ g/cm}^3 \times 5 \text{ cm}^3$. $m(\text{silver}) =$

52.5 gm $m(\text{gold}) = 96.5 \text{ g}$. Answer

gold. 6. Five mL of ethanol has a mass of 3.9 g, and 5.0 mL of benzene has a mass of 44 g. Which liquid is denser?

$v(\text{ethanol}) = 5 \text{ ml}$ $v(\text{benzene}) = 5 \text{ ml}$.

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