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emission of electrons from a metal's surface when light of a certain frequency shines on it 9. A figure indicating the relative sizes and energies of atomic orbitals Describe how each pair is related. 10.

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central atom with two identical bonds
to other atoms and two lone pairs.
The shape of this ...

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are not in a straight line. This geometry occurs when the central atoms contain 4 electron groups (2 bonding and 2 nonbonding) or 3 electron groups (2 bonding and 1 nonbonding) The bond that results when two nonmetals combine in a chemical reaction.

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Electrons move in circular orbits .
around . the nucleus . at . fixed.
energy. levels. Electrons are never
between energy levels or energy
shells. An electron must have . just the

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right amount of energy. to jump from
one level to another. A Chapter 13
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The electrons in the outer orbit
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PLAY. 3×10^8 m/s. What is the

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speed of light? meters. What unit
must wavelength be in? $E=hv$.
Equation to solve for energy. $v= e/h$.
Equation for frequency. Joules. In
what units is energy measured. 6.626
 $\times 10^{-34}$ JxS. Planck's constant. Speed
of light formula. $c = \quad \times v$.

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Chapter 5: Electrons in Atoms Models of the Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun.

Rutherford ' s model fails to explain

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why objects change color when
heated.

Chapter 5 Electrons In Atoms Answer
Key

Chapter 5: Electrons in Atoms 5.1
Wave-Particle

Duality/Electromagnetic

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Spectrum/Relationship of
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Valence Electrons

Chapter 5: Electrons in Atoms

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Chapter 9: Electrons in Atoms and the
Periodic Table Ch 9 Page 1 . Visible
light -the form of electromagnetic
radiation that we can "see"
Microwaves, radiowaves, x -rays,
ultraviolet radiation, infrared radiation
-the forms of electromagnetic

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radiation that we can't see.

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CHAPTER 5 Electrons in Atoms

Resource Manager Objectives Section

Section 5.1 Light and You will express
the arrangements of electrons in

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atoms through orbital notations,
electron Analysis Answers will vary.
Students will determine that
observations typically rely heavily
upon sight, ...

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Although the speed of all electromagnetic waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in

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other words, as one quantity
increases, the ...

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