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Electron Configuration Make a Foldable to help you summarize the three rules that define how electrons are arranged in an atom.
STEP 1 Fold a sheet of paper in half lengthwise. Make the back edge about 2 cm

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longer than the front edge. STEP 2 Fold into thirds.

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Arrangement of Electrons in Atoms SECTION 3
SHORT ANSWER Answer the following questions in the space provided. 1. State the Pauli

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exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states. The Pauli exclusion principle states that no two electrons in an atom may have the same set of four quantum numbers.

~~4 Arrangement of Electrons in Atoms~~

Chapter 18 Study Guide 1. Why do atoms gain or lose electrons? To get a stable _electron ___ _configuration ___. (full outer shell - octet) 2. Which group on the periodic table contains elements with stable electron configurations? ___ noble ___ ___gases _____

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Group 18 3. If you have a full outer energy level you have a stable __octet __ or __configuration __. 4.

~~Key_Chapter_18__compounds_Study_Guide~~
~~Chapter 18 Study ...~~

Chemistry Chap. 5.3 Study Guide (Electron configuration) ?questionThe arrangement of electrons in an atom is called the atom's _____. answerElectron configuration questionElectrons in an atom

~~Chemistry Chap. 5.3 Study Guide (Electron configuration ...~~

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Each orbital may contain at most (two, four) electrons. 16. All s orbitals are (spherically shaped, dumbbell shaped). . A principal energy has (n, '12) energy sublevels. The maximum number of (electrons, orbitals) related to each principal energy level equals $2n^2$. There are (three, five) equal energy p orbitals.

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You can use the atomic number & atomic mass to help you find the number of protons, electrons, & neutrons: # of protons - use the atomic number # of electrons - is always the

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same as the number of protons, so you can also use the atomic number # of neutrons - round the atomic mass, and then subtract atomic mass - atomic number The number of protons & electrons will always be the same, but the number of neutrons can be different.

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Each orbital may contain at most (two, four) electrons. All s orbitals are (spherically shaped, dumbbell shaped). A principal energy has (n, 112) energy sublevels. The maximum number of (electrons, orbitals) related to

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each principal energy level equals $2n^2$. There are (three, five) equal energy p orbitals.

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The three parts of an atom are protons, neutrons, and electrons. Protons carry a positive electrical charge. Neutrons are electrically neutral. Electrons carry a negative charge, equal in magnitude to that of a proton. Protons and neutrons stick together to form the atomic nucleus. Electrons orbit around the nucleus.

~~Basic Atomic Structure and Atomic Theory~~

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Atoms can form chemical bonds by sharing electrons. Atoms with more than 3 and less than 6 electrons in the outer shell tend to form chemical bonds by sharing electrons. Consider the reaction between hydrogen and carbon. STUDENT STUDY GUIDE FOR 8TH GRADE CHEMISTRY

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Chapter 5: Electrons in Atoms Models of the Atom Rutherford used existing ideas about the

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atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun.

~~Chapter 5 Electrons In Atoms Study Guide Answers~~

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charged particle in an atom's nucleus is the _____. An element's identity can be determined from

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Electrons. are particles that carry a .

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negative charge. and . surround. the . nucleus. in what is know as the . electron cloud on energy levels. There is a limit to the number of electrons each energy level can hold. The first energy level holds 2 electrons, the second holds 8, the third holds 18, the fourth holds 32, and the fifth holds 50. The . number. of . electrons

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