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Academy Empirical Formula

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~~Respiration: Crash Course~~

*~~Biology #7~~ **Empirical Formula***

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The empirical formula for a

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Compound is C_2H_5 and its relative formula mass is 58.0. Deduce its molecular formula. Deduce its molecular formula. (Relative atomic masses: $C = 12.0$, $H = 1.0$)

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is CH_2Cl , mass is 99.0

g/mole. c. Empirical formula
is C_3H_4 mass is 121 g/ mole.

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empirical formula. 2. Divide

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the molecular mass by the empirical mass. 3. Multiply all of the subscripts in the empirical formula by the resulting factor. Example Problem: What is the molecular formula for a material that is 92.2% C and

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7.80% H by mass? The
material has a molar mass of
52.0 g/mol. 1. Molecular
mass /

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