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*CHEM113L:  
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*Post-lab Analysis*

## **Post Lab: Determination of an Equilibrium Constant**

*Determining an  
Equilibrium Constant  
by Spectrophotometry  
Procedure 105*

## **Equilibrium 5 Finding A Constant Post Lab Lab**

~~Experiment #13: The  
Equilibrium Constant.~~

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~~Calculation of Post  
Equilibrium Constant  
Lab Equilibrium~~

*Constant Lab Part 1:  
K, Beer's Law, and  
Stoichiometry*

*Determination of  $K_{eq}$   
for  $FeSCN^{2+}$  Lab  
Explanation Video*

*Lab Experiment #13:  
Equilibrium Constant*

*Equilibrium Constant  
Prelab Video McMaster  
Chem 1A03 / 1E03*

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*Experiment 3 -  
Determination of an  
Equilibrium Constant  
How To Calculate The  
Equilibrium Constant  
K - Chemical  
Equilibrium Problems  
Ice Tables  
Solubility Virtual Lab  
Equilibrium Constant  
Kc - table method  
Equilibrium  
Equations: Crash  
Course Chemistry*

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## #29 Determination of an Equilibrium Constant - WJEC A

### Level Experiment

Practice Problem:

Calculating

Equilibrium

Concentrations Le

Chatelier's Principle

Demonstration

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Ice Table -

Equilibrium Constant

Expression, Initial

Concentration,  $K_p$ ,

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Kc, Chemistry Post

Examples Limiting

Reactant Experiment

~~General lab 106 and~~

~~109~~ **Calculating**

**Equilibrium**

**Concentrations-1**

~~The Equilibrium~~

~~Constant and~~

~~Expression with~~

~~Examples~~

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Spectrophotometric

determination of an

equilibrium constant



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~~Equilibrium Constant  
calculations~~

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~~Keq FeSCN<sub>2</sub><sup>+</sup> Lab  
Equilibrium Constant~~

**The Determination  
of an Equilibrium  
Constant CHEM**

~~1146: Determination  
of an Equilibrium~~

~~Constant CHM 116–  
Determination of~~

~~Equilibrium Constant  
Lab Lecture Theory~~

~~and Calculations~~

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ALEKS - Calculating  
Equilibrium

Composition from an  
Equilibrium Constant

~~16.6h Using the~~

~~general properties of  
equilibrium constants~~

~~Equilibrium Constant~~

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equilibrium constant can then be determined for each mixture; the average should be the equilibrium constant value for the formation of the  $\text{FeSCN}^{2+}$  ion. In Part A of this experiment, you will prepare  $\text{FeSCN}^{2+}$  solutions of known concentrations,

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chemistry 2 lab  
experiments in.

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motion friction.

~~Equilibrium Constant  
Post Lab Answers~~  
Answer to Experiment  
6: DETERMINATION  
OF AN EQUILIBRIUM  
CONSTANT Post-Lab



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Questions 1. To  
determine the  
accuracy of your  
average e...

~~Solved: Experiment 6:  
DETERMINATION  
OF AN ...  
chegg.com~~

This video is about  
the AP Chemistry Lab  
Experiment #13: A  
Spectrometric  
Determination of  $K_{eq}$

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Equilibrium  
of the Constant Post  
Iron(III)-Thiocyanate  
Lab Answers  
System. In this video  
you will lea...

~~Lab Experiment #13:  
The Equilibrium  
Constant. YouTube~~

The equilibrium  
constant,  $K$ , is used to  
quantify the  
equilibrium state. The  
expression for the  
equilibrium constant

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for a reaction is determined by examining the balanced chemical equation.

## ~~Experiment 3 Determination of an Equilibrium Constant for ...~~

The experiment was completed to determine the equilibrium constant

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of a chemical reaction

This chemical reaction created a state of chemical equilibrium. than one, or small, less than one, but not negative or zero. A large  $K_c$  (equilibrium constant) means that at equilibrium, the concentrations of the products will generally be greater than the

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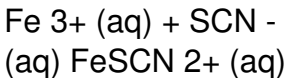
~~Experiment 34: Lab  
Report About The  
Equilibrium Constant~~

: Calculate the concentration of  $\text{SCN}^-$  at equilibrium for Trials 1-4 using the equation:  $[\text{SCN}^-]_{\text{eq}} = [\text{SCN}^-]_{\text{i}} - [\text{FeSCN}_2^+]_{\text{eq}}$  7. Calculate  $K_c$  for Trials 1-4. Be sure to show the  $K_c$  expression and the

# Acces PDF Equilibrium

values substituted in for each of these calculations. 8. Using your four calculated  $K_c$  values, determine an average value for  $K_c$ . How constant were your  $K_c$

~~Lab 1: Chemical  
Equilibrium: Finding a  
Constant,  $K_c$~~



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The equilibrium constant for equation 4 can be expressed using the concentrations of the three components: ( 5 )  $K = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^-]}$  In order to calculate the equilibrium constant, one must simultaneously determine the concentrations of all

Acces PDF  
Equilibrium  
three of the Post  
components.  
Lab Answers

~~Lab 5 – Determination  
of an Equilibrium  
Constant~~

Question: Pre-Lab –  
The Determination Of  
An Equilibrium  
Constant DATA

TABLE Parts I And II  
Beaker [FeSCN<sup>2+</sup>]  
Absorbance 1 0.00 M  
0.001 2 8.0 X 10<sup>-5</sup> M



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0.238 3  $1.2 \times 10^{-4}$  M

0.344 4  $1.6 \times 10^{-4}$  M

0.435 5  $2.0 \times 10^{-4}$  M

0.530 Unknown, Part  
II 0.302 Insert Excel

Absorbance Vs

Concentration Curve

Here Include Linear

Least Squares Fit Too

1. (Part II) Use The  
Calibration ...

~~Solved: Pre-Lab~~

~~The Determination Of~~

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~~An Equilibrium Cons~~

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Folake Oladele CHM  
102 7/21/17  
Experiment 34 Post-  
lab Questions 1. The  
experiment I would

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Conduct is to measure the absorbance of  $\text{FeNCS}_2^+$  at various ranges, starting from 300nm to 500nm, to analyze UV-visible spectra. From that graph, if 447 nm is the optimal wavelength, the absorbance at that particular wavelength should be higher than the other.

2. A) Since

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Constant is  
proportional to ...  
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...

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thiocyanate~~

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100%. = 10 ??bc = 10  
?A. Then, you will use  
your understanding of  
equilibrium processes  
to deduce the  
equilibrium  
concentrations of the  
reactants. Knowing all  
three concentrations  
listed above in Eq. 2.

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$K = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^-]}$  allows the equilibrium constant for this reaction to be calculated.

~~Lab 11~~

~~Spectroscopic  
Determination of an  
Equilibrium ...~~

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Science laboratory  
reports are developed  
to interact the findings  
of research, in such a  
way that is clear to  
readers. You need to  
not forget to include  
any extra details,

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which may be helpful  
for readers.

~~Spectrophotometric  
Determination Of An  
Equilibrium ...~~

Question: Solubility  
Product Constant  
Post-Lab Questions  
And Exerciss All  
Questions Must Be  
Answered During The  
Lab And Submitted  
With Your Lab

# Acces PDF Equilibrium

Report) Please  
Answer The Following  
Questions And Show  
All Work And Units  
Q1.

## ~~Solubility Product Lab~~ Answers

The lab handout is attached below in both Word and PDF format. (Note: Hat tip to Mr. Tony Locke, who shared this lab

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with an IB workshop I attended. I have modified it some along the way.) The lab is a pretty basic look at causing disruptions to an equilibrium and making predictions - and then observations - based on Le Châtelier's Principle.

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