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actually writing the
phrase "thermodynamic

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equilibrium". Referring to systems closed to exchange of matter, Buchdahl writes: "If a system is in a terminal condition which is properly static, it will be said to be in equilibrium ."

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The goal of this experiment is to determine the value of an equilibrium constant at different temperatures and use these data to calculate the enthalpy and entropy of reaction.

The value of an equilibrium constant for a reaction varies, depending on the temperature. In

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endothermic reactions,
the value of K increases
as the ...

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thermodynamics often
find it to be one of the
most challenging
subjects, especially at
more advanced levels
where calculus is an

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integral part of the topic. It is, however, an important topic in both Physics and Chemistry and is something that all physical scientists, engineers, and even many biologists study.

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the concept of
temperature from the
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Use (2.255) and (2.205)
to obtain the relation

$$T - S = T V C P$$

2. (2.257) Suggestions
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J. Adkins, Equilibrium
Thermodynamics, third
edition, Cambridge
University Press (1983).
Ralph Baierlein,
Thermal Physics,
Cambridge University
Press, New York (1999).
Robert P. Bauman,

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Statistical Mechanics,
Macmillan Publishing
Company (1992).

255 and 2205 to obtain
the relation $\frac{1}{T} = \left(\frac{\partial S}{\partial U}\right)_V$...

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the systems $\text{CO}_2 +$
 $\text{NH}_3 + \text{H}_2\text{O}$, $\text{CO}_2 +$
 $\text{H}_2\text{S} + \text{H}_2\text{O}$, $\text{H}_2\text{S} +$
 $\text{NH}_3 + \text{H}_2\text{O}$, and CO_2
 $+ \text{NH}_3 + \text{H}_2\text{S} + \text{H}_2\text{O}$.

The types of data in this
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types of equilibrium data, including both equilibria in solution and vapor-liquid equilibrium data, enthalpies, heat capacities, and densities. There are 215 references cited.

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