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An acid is a compound that donates a proton. After the proton leaves, the remaining compound (with one less proton) is the conjugate base. Example: $\text{NH}_4^+ + \text{H}_2\text{O} \rightarrow \text{NH}_3 + \text{H}_3\text{O}^+$ Acid = NH_4^+ Conjugate base = NH_3

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Holt Chemistry Acid And Bases Test Answers reactants is the Brønsted-Lowry acid, and which is the Brønsted-Lowry base? Explain your answer. $\text{H}_2\text{CO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{HCO}_3^-(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$. The H_2CO_3 is the Brønsted-Lowry acid because it donates a proton to the H_2O . The H_2O is the Brønsted-Lowry base because it accepts the proton. b.

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