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Chemical equilibrium is a dynamic process consisting of forward and reverse reactions that proceed at equal rates. At equilibrium, the composition of the system no longer changes with time. The composition of an equilibrium mixture is independent of the direction from which equilibrium is approached.

The Concept of Chemical Equilibrium

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Chemical equilibrium: A state in which the rates of the forward and reverse reactions are equal and the concentrations of the reactants and products remain constant. ? Equilibrium is a dynamic process @ the conversions of reactants to products and products to reactants are still going on, although there is no net change in the number of reactant and product molecules.

Chapter 14. CHEMICAL EQUILIBRIUM

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Holt Chemistry Concept Review Answers Chemical Equilibrium When a chemical reaction and it's reverse occur at the same time and rate the reaction is said to be in Equilibrium According to la chatelier principle

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Chemical Equilibrium | Holt: Modern Chemistry | N... a reversible reaction in which there is no longer any change in the concentration of reactants and products. chemical equilibrium. a state of balance in which the rate of a forward reaction equals the rate of the reverse reaction and the concentrations of products and reactants remain unchanged. equilibrium. Holt Chemistry Chapter 14 Chemical

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Cato Maximilian Guldberg (1836-1902), and Peter Waage (1833-1900) Law of mass action – relationship between concentrations of reactants and products at equilibrium. If aA + bB ? pP + qQ, then an equilibrium expression can be constructed. Kc = [P]^p[Q]^q / [A]^a[B]^b. equilibrium expression depends only on stoichiometry of reaction and not mechanisms. equilibrium constant:

15.5: Chemical Equilibrium (Summary) — **Chemistry LibreTexts**

Our most basic concept of equilibrium is based on the observation that change in an isolated system eventually ceases; once change ceases, it never resumes. In this book, we call the idea of a static state of an isolated system the primitive equilibrium. We also observe that change eventually ceases in a closed system that is not isolated but whose temperature, pressure, and volume are kept constant.

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