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Determine the molecular weight and stoichiometric mole and mass air-fuel ratios for the Oklahoma gas mole composition given in Table 3.4. Solution: Equation (3.2), CH4+ 2O2+ 2(3.76)N2CO2+ 2H2O + 2(3.76)N2(3.2) shows that there are 2 + 2(3.76) = 9.52 moles of air required for complete combustion of each mole of methane.

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This solutions manual has been prepared to accompany the 3rd edition of the author's Introduction to Internal Combustion Engines. At the end of many of the questions is a discussion, which is intended to provide useful supplementary information.

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Apr 26, 2018 - Solutions Manual for An Introduction to Combustion Concepts and Applications 3rd Edition by Stephen R. Turns 0073380199 9780073380193

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