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0.25 g 25 cg c. 400 cm<sup>3</sup> 0.4 L d. 400 cm 3 0.0004 m 2. For each measuring device shown below, identify the quantity measured and tell when it would remain constant ...

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2(g) ? NO<sub>2</sub>(g) a. Balance the equation. 2N<sub>2</sub>O(g) + 3O<sub>2</sub>(g) ? 4NO<sub>2</sub>(g) 4 mol NO<sub>2</sub>:3 mol O<sub>2</sub> b. What is the mole ratio of NO<sub>2</sub> to O<sub>2</sub>? 15.0 mol c. If 20.0 mol of NO<sub>2</sub> form, how many moles of O<sub>2</sub> must have been consumed? True d. Twice as many moles of NO<sub>2</sub> form as moles of N<sub>2</sub>O are consumed. True or False? False e. Twice as many grams of NO<sub>2</sub> ...

mc06se cFMsr i-vi

CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

5 The Periodic Law

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SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms.

6 Chemical Bonding

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