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HOMEWORK 9-2 (pp. 280-282)
VOCABULARY ... Circle the letter of the
best answer. 1. Given the balanced
equation $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 +$

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10H₂O, how many moles of CO₂ are produced when 14.9 g of O₂ ... Microsoft Word - Homework 09-01 Author: Susan M. Free

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CHAPTER 9 REVIEW Stoichiometry
SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88%
The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: N₂(g) + 3H₂(g) → 2NH₃(g)

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CHAPTER 6 REVIEW Chemical Bonding
SECTION 3 SHORT ANSWER Answer the following questions in the space

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provided. 1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b ...

~~6-Chemical Bonding~~

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Stoichiometry. Sections 1-3 Introduction to Stoichiometry Ideal. Stoichiometric Calculations Limiting Reactant and. Percent Yield. 2. Definitions. Composition stoichiometry deals with the mass. relationships of elements in compounds. Reaction stoichiometry involves the mass.

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