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Shape and Bond Angles
(Methanol) Predict the
Molecular Geometry of
I₃⁻ (triiodide ion)

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#3

Molecular Geometry

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Molecular Geometry 3

9. Explain the difference

between a bonding

electron domain and a

nonbonding electron

domain using the

examples in Model 1.

10. Circle the correct

word or phrase to

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Complete the sentences:

Pairs of electrons will
(attract / repel) each
other.

20 Molecular Geometry-
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community ensures ...

Molecular Geometry

How can molecular

shapes be predicted

using the VSEPR

theory? why? When you

draw a Lewis structure

for a molecule on paper,

you are making a two-

dimensional representa-

tion of the atoms. In

reality however,

molecules are not flat-

they are three-

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dimensional. The 3D
shape of a molecule is
important because it
determines many
physical and chemical
properties for the
substance

can shapes be rhe
theory? whv? - LTHS
Answers

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The valence shell
electron pair repulsion
(VSEPR) theory (or
"VESPER" for short) is
how the geometry of a
molecule is determined

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Answer Key
around a central atom.
The molecular
geometry main shapes
are tetrahedral, trigonal
planar, trigonal
pyramidal, bent, and
linear and are named by
measuring the bond
angles between the
central atom and
another atom...

Molecular Geometry

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Vsepr Theory Pogil

Worksheet Answers

13, Based on the

information in the Read

This! section, sketch one

of the molecular shapes

shown below in each of

the boxes provided in

Model I. Trigonal

planar 120° Bent

104.5° Three-

Dimensional Molecular

Shapes Tetrahedral

109.5° STOp

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Linear Pyramidal 1070
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