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must be

considered: CuO

$(s) + \text{H}_2 (g) \rightarrow$

$\text{H}_2\text{O} (l) + \text{Cu}$ at

equilibrium α - β

$n\text{O} \text{D} \text{ } 2 - \alpha \text{ Cu} (s)$

$+ \frac{1}{2} \text{O}_2 (g) \rightarrow \text{CuO}$

1 at equilibrium

α - β $n\text{O} \text{D} \text{ } 2 - \beta \text{ } 2$

In the final

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state, only O_2 is present.
Therefore $\alpha = nOD_2$. In an excess of O_2 , all the 1 copper is oxidized.
Therefore $\alpha - \beta = 0$ or $\beta = nOD_2$.

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