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Processes (Ideal Gas) A steady flow compressor handles $113.3 \text{ m}^3 / \text{min}$ of

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nitrogen ($M = 28$; $k = 1.399$) measured at intake where $P_1 = 97$ KPa and $T_1 = 27$ C. Discharge is at 311 KPa. The changes in KE and PE are negligible. For each of the following

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2 3 energy J N m kg m power = = = =
time s s s charge current = time
charge = current*time = A s energy

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power = current * electric potential
time 2 3 energy kg m electrical
potential = current * time A s electrical
potential current = resistance 2 23

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The reaction is $\text{H}_2\text{O} + \text{CO} = \text{H}_2 + \text{CO}_2$

Gf. Chapter 14 Practice Problem

Solutions To accompany Introductory Chemical Engineering

Thermodynamics. J.R. Elliott, C.T.

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(02/11/02) 16 (P14.1) An equimolar mixture of H_2 and CO can be obtained by the reaction of steam with coal.

Compute the equilibrium compositions at 550 C based on an equimolar feed of H_2 , CO , and H_2O .

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H2O. The reaction is ...*

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03 A keen student of thermodynamics was thinking for a while about the true meaning of specific heat capacity, C_p . There seem to be many versions, with various symbols used in the literature.

a) For a pure substance, how many independent variables are needed to define C_p .

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Thermodynamic Properties

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PROPERTIES AND OTHER BASIC
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AND OTHER BASIC CONCEPTS 1.1
PRELIMINARY CONCEPTS – THE
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In order to accurately and precisely discuss various aspects of thermodynamics, it is essential to have a well-defined vernacular. As such, a list of some foundational concepts and their definitions are shown

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these of thermodynamic relations in the practical attack on chemical problems, with a willing acceptance of arithmetic, graphs, and rough estimates. These characteristics are retained in the revision. With regard to the first, we agree reluctantly that the revisors were wise to keep the unsystematic and ad hoc derivations, even though the typical ...

Thermodynamics (Lewis, G. N.; Randall, M.)

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terms of microscopic constituents by .!

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Chemical thermodynamics is the study of the interrelation of heat and work with chemical reactions or with physical changes of state within the confines of the laws of thermodynamics. Chemical thermodynamics involves not only laboratory measurements of various thermodynamic properties, but also the application of mathematical methods to the study of chemical questions and the spontaneity of ...

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