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Objectives. After completing this section, you should be able to. write the equilibrium constant expression for a given reaction. assess, qualitatively, how far a reaction will proceed in a given direction, given the value of K_{eq} ; explain the difference between rate and equilibrium.

6.8: Describing a Reaction - Equilibria, Rates, and Energy ...

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Value of K_{eq} ; explain the difference between rate and equilibrium

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developed as a correlation of reactivities (rates) and equilibria in reactions of meta- and para-substituted benzene derivatives. As early as 1912 it was noted that in reactions of m- and p-substituted benzene

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derivatives there were regularities in substituent effects. For example, whatever effect a p-NO₂ group

CHAPTER 1 SUBSTITUENT EFFECTS ON ORGANIC RATES AND EQUILIBRIA

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In any reversible reaction such as $\dots v_A A + v_B B \leftrightarrow v_R R + v_S S$ [2.1] \dots the system inevitably moves toward a state of equilibrium, or maximum probability. This equilibrium state is very important in analyzing chemical reactions because it defines the limit to which any reaction can proceed. Organic reactions, particularly those constituting a synthetic scheme for a fine chemical ...

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5.1 Rates, equilibrium and pH. 5.1.2 How far? fi) the qualitative effect on equilibrium constants of changing temperature for exothermic and endothermic reactions; fii) the constancy of equilibrium constants with changes in concentration, pressure or in the presence of a catalyst

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