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It has one unpaired electron in the 3s orbital. The 3s valence atomic orbital of

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sodium overlaps with another such orbital of the same energy to form molecular orbitals. The atomic orbitals continue to combine in some fashion forming a band.

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 $\text{Fe (s)} + \text{FeSO}_4 \text{ (aq)} \rightarrow \text{no reaction.}$
 $\text{Fe (s)} + \text{Al}_2 \text{ (SO}_4)_3 \text{ (aq)} \rightarrow \text{no reaction.}$
 $\text{Fe (s)} + \text{ZnSO}_4 \text{ (aq)} \rightarrow \text{no reaction.}$
No reaction occurs when any of the metals from zinc, iron, aluminium, and copper are placed in aqueous solution of $\text{Al}_2 \text{ (SO}_4)_3$.

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