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of each of the following
compounds? H_2O CH_3OH $(\text{NH}_4)_2\text{CO}_3$
 C_2H_8 2. Empirical and
Molecular Formulas Give the
empirical formula for each
compound: N_2H_2 NH_3
 BH_3 BH_3 H_2O_2
 OH ...

Stoichiometry Compositions

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Stoichiometry. The atomic ratios in each compound are also the relative number of atomic mass units of its elements. The first example is nitrous oxide (N_2O), as shown in Table 1. The relative masses were obtained by multiplying the atomic ratios and atomic masses. ... CliffsNotes study guides are written by real teachers and professors, so ...

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regarded as unclean the
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stoichiometry and the
balanced equation. For each
statement below, write true
or false. _____true_____ 1.
The study of the

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quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

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What is stoichiometry? 1.
true 2. true 3. false 4.
true 5. true 6. 2, 2, 64.10
7. 3, 3, 96.00 8. 2, 2,
88.02 9. 4, 4, 72.08 10.
methanol and oxygen gas 11.
carbon dioxide and water 12.
160.10 g 13. 160.10 g 14.
They are equal. 15. A mole

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